Write the following information in the first page of Answer Script before starting answer

ODD SEMESTER EXAMINATION: 2020-21

Exam ID Number	
Course	Semester
Paper Code	Paper Title
Type of Exam:	(Regular/Back/Improvement)

Important Instruction for students:

- 1. Student should write objective and descriptive answer on plain white paper.
- 2. Give page number in each page starting from 1st page.
- 3. After completion of examination, Scan all pages, convert into a single PDF, rename the file with Class Roll No. **(2019MBA15)** and upload to the Google classroom as attachment.
- 4. Exam timing from 10am 1pm (for morning shift).
- 5. Question Paper will be uploaded before 10 mins from the schedule time.
- 6. Additional 20 mins time will be given for scanning and uploading the single PDF file.
- 7. Student will be marked as ABSENT if failed to upload the PDF answer script due to any reason.

REV-01 BPH

B.PHARM. FIRST SEMESTER PHARMACEUTICAL ANALYSIS-I BP-102 T

Duration : 3 hrs.

Time : 20 min.

Full Marks: 75

(<u>PART-A : Objective</u>)

Marks : 20 1X20=20

Choose the correct answer from the following: 1. API stands for?

1.	a. Active Pharmacy Interpretation c. Active Pharmaceutical ion	b. Acute Pulmonary Interaction d. Active Pharmaceutical ingredients
2.	Which of the following is widely used for qua a . Non Aqueous titration c . Acid base titrations	antitative analysis? b. Aqueous titrations d. Volumetric Analysis
3.	 % w/v express: a. Number of ml of solute in 100 gm of product c. Number of gm of solute in 100 ml of product 	b. Number of ml of solute in 100 ml of productd. None of these
4.	Random error is also known as: a. Accidental error c. Determinate error	b. Indeterminate error d. Both a and b
5.	Following are the types of systemic error exce a . Errors of methods c . Personal error	ept: b. Instrument method d. Random
6.	According to Bronsted theory a. Acid c. Base	is a substance that can accept protons. b. Buffer solution d. Both a and b
7.	Amphiprotic solvents are both: a. Aprotic, Protophilic c. Protogenic, Aprotic	b. Protophilic, Protogenic d. None of these
8.	Non Aqueous titrations are based on: a. Arrhenius, theory c. Bronsted Lowry theory	b. Lewis theoryd. None of this
9.	Which of the following is not an indicator of a. Naptholbenzoin c. Crystal violet	non aqueous titration? b. Oracet blue B d. Complexometric titartion

10. Which method is based on precipitation of aa. Fajan's methodc. Volhard's method	dsorption indicators? b. Mohr's method d. None of these
11. Which method is used for Water Analysis?a. Naptholbenzoinc. Crystal violet	b. Oracet blue B d. Complexometric titartion
12. Which among the following can act as a maswith EDTA?a. Thioglycerola. Botassium guanida	sking agent for complexometric titration b. Aluminium floride d. Triethanolamine
 c. Potassium cyanide 13. Stability complex is denoted by: a. L c. K 	b. M d. N
14. Which of the following is not a property of aa. Easily filtered and washed free of contaminantsc. Unreactive with constituents of the atmosphere	a good precipitate? b. Significant loss of analyte during filtration d. Both a and b
 15. Inthe analyte is separated fractional is converted to a compound of known of a. Volatilization gravimetry c. Electrogravity 	
 16. Aromatic amino group is reacted with diazonium salts. a. Sodium peroxide c. Silver nitrite 	 in cold acid solution to form b. Sodium nitrite d. Sodium hydroxide
 17. Compounds possess	·
18. Which of the following is not an Oxidizing aa. Hydrogen Peroxidec. Nitric acid	ngent? b. Sulphuric acid d. Formic acid
19. Which of the following is not a reducing agea. Alkali earth metalc. Peroxy disulfuric acid	nt? b. Formic acid d. Sulphite compounds
20. SI units of conductance:a. mhoc. volt	b. siemens d. None of these

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(<u>PART-B : Descriptive</u>)

Tiı	Time : 2 hrs. 40 min.				
	[Answer any seven (7) questions]				
1.	Write down the applications of Conductometry.	5			
2.	Explain the theory of Redox titrations.	5			
3.	Write down the applications of diazotization titrations.	5			
4.	What is colloidal state? Enumerate the properties of colloidal particles.	5			
5.	Write down the various types of complexometric titrations.	5			
6.	Define Systemic error and explain in brief the types of systemic error.	5			
7.	List out some indicators used in non aqueous titration.	5			
8.	Write about the estimation of Sodium Chloride (Mohr's method).	5			
9.	Explain about Volhard's method.	5			

(<u>PART-C : Long type questions</u>)

[Answer any two (2) questions]

1.	Explain principle, assay procedure and types of diazotization titrations.	10
2.	Discuss the neutralization curve for strong acid v/s strong base titration and show how pH changes during the titration.	10
3.	Explain in details about the acid base titrations.	10

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